



A Publication  
of Reliable Methods  
for the Preparation  
of Organic Compounds

## Working with Hazardous Chemicals

The procedures in *Organic Syntheses* are intended for use only by persons with proper training in experimental organic chemistry. All hazardous materials should be handled using the standard procedures for work with chemicals described in references such as "Prudent Practices in the Laboratory" (The National Academies Press, Washington, D.C., 2011; the full text can be accessed free of charge at [http://www.nap.edu/catalog.php?record\\_id=12654](http://www.nap.edu/catalog.php?record_id=12654)). All chemical waste should be disposed of in accordance with local regulations. For general guidelines for the management of chemical waste, see Chapter 8 of Prudent Practices.

In some articles in *Organic Syntheses*, chemical-specific hazards are highlighted in red "Caution Notes" within a procedure. It is important to recognize that the absence of a caution note does not imply that no significant hazards are associated with the chemicals involved in that procedure. Prior to performing a reaction, a thorough risk assessment should be carried out that includes a review of the potential hazards associated with each chemical and experimental operation on the scale that is planned for the procedure. Guidelines for carrying out a risk assessment and for analyzing the hazards associated with chemicals can be found in Chapter 4 of Prudent Practices.

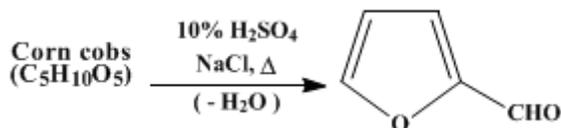
The procedures described in *Organic Syntheses* are provided as published and are conducted at one's own risk. *Organic Syntheses, Inc.*, its Editors, and its Board of Directors do not warrant or guarantee the safety of individuals using these procedures and hereby disclaim any liability for any injuries or damages claimed to have resulted from or related in any way to the procedures herein.

*These paragraphs were added in September 2014. The statements above do not supersede any specific hazard caution notes and safety instructions included in the procedure.*

*Organic Syntheses, Coll. Vol. 1, p.280 (1941); Vol. 1, p.49 (1921).*

## FURFURAL

### [2-Furaldehyde]



Submitted by Roger Adams and V. Voorhees.

Checked by H. T. Clarke and E. R. Taylor.

### 1. Procedure

In a 12-l. round-bottomed flask are placed 1.5 kg. of dry corn cobs (ground to about the size of corn kernels) (Note 1), 5 l. of 10 per cent sulfuric acid, and 2 kg. of salt. The flask is shaken in order to secure a homogeneous mixture and is then connected with an upright tube, water condenser, and return tube as shown in Fig. 17, p. 282 (Note 2). Heat is applied from a ring burner, the flame being adjusted so that the liquid distils at a rapid rate.

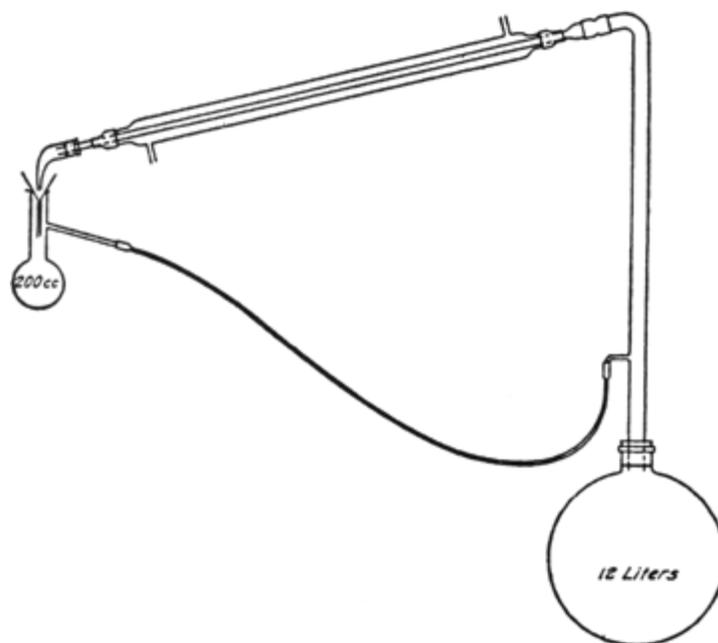
The distillation process is continued until practically no more furfural can be seen collecting in the distilling flask used as a receiver. The above operation requires from five to ten hours.

This distillate is now treated with enough sodium hydroxide so that the mixture is left just faintly acid, and the furfural separated. It amounts to 180–220 g. The wet furfural is distilled under reduced pressure from a Claisen flask which is heated in an oil bath (Note 3). The temperature of the bath is never permitted to rise above 130°. At first, water together with some furfural distils, and this fraction is separated to be worked up with a later portion. Finally, 165–200 g. of pure furfural (b.p. 90°/65 mm.; 159°/745 mm.) distils, and this fraction, collected separately, is found to be practically colorless. The distillation of the crude material must not be carried out under ordinary pressure; otherwise the product turns dark rapidly on standing. After one distillation under reduced pressure as described, however, a distillation under atmospheric pressure may be carried out without the objectionable results just mentioned (Note 4).

### 2. Notes

1. Certain samples of ground cobs which had remained in the laboratory for a year did not give nearly such good yields of furfural as fresher ones.
2. The apparatus described (Fig. 17) is very convenient for laboratory use, but may be modified in many ways as long as a few essential points are kept in mind. A modification worth mentioning is the use as a receiver of a vessel the bottom of which has a stopcock attached, thus allowing the furfural to be drawn off at any time. It is necessary to have the receiver several inches higher than the opening in the upright tube so that the aqueous liquors will flow back to the reaction flask; the lower end of the funnel in the receiver must be below the side arm of the receiver in order to prevent bubbles of furfural from collecting on the surface of the liquid and being carried back to the reaction flask; the entrance of the returning liquid into the upright tube must be at such a point that a vigorous stream of vapor passes through the returning liquid, thus extracting much of the furfural carried back in the water; the tube for returning the aqueous liquors must at some point be lower than the entrance to the upright tube so that a trap of liquid will be formed and prevent the vapors from the reaction flask entering the side tube. If the return tube is small and swings too low, it happens occasionally that a bubble of air gets into the tube and prevents the regular flow of liquid; this can be remedied by shaking the tube until the difficulty is overcome. It was found that the more efficient the upright tube was as a fractionating column, the more complete was the separation of furfural.

Fig. 17.



3. Distillation of the product under reduced pressure is essential. Moreover, in this final distillation the precautions mentioned (the use of an oil bath and an outside temperature of less than  $130^{\circ}$ ) must be carefully observed. When the [furfural](#) is distilled under ordinary pressure or when it is distilled under reduced pressure with a free flame, a practically colorless product is at first obtained. After a few days and sometimes after a few hours this product will gradually darken until finally a black liquid results. This change, although most marked in the presence of light, takes place readily even when the aldehyde is stored in brown glass bottles. On the other hand, when the crude [furfural](#) is distilled under reduced pressure and is never heated above  $130^{\circ}$  during the process of distillation, a product is obtained that develops only a slight color when exposed to direct sunlight during several days. Further purification according to methods described in the literature does not give a product which will remain colorless on standing.

4. The low yield of [furfural](#) from straw and other materials makes it desirable to extract the aqueous [furfural](#) distillate after acid hydrolysis by a solvent heavier than water, such as [chloroform](#). This is done by adding a funnel to the end of the condenser so that the distillate passes directly into a layer of [chloroform](#) contained in a receiver. Distillation is continued until the [chloroform](#) layer does not increase. The [furfural](#) is obtained by distillation of the [chloroform](#) layer, and practically all the [chloroform](#) is recovered (H. I. Waterman, private communication).

### 3. Discussion

Practically quantitative yields of [furfural](#) are obtained when pentoses are subjected to the action of [hydrochloric acid](#).<sup>1</sup> Carbohydrate materials such as corn cobs, wood, the hulls of oats, rice, peanuts, etc., when heated with steam under pressure or distilled with dilute hydrochloric or sulfuric acids, yield appreciable quantities of [furfural](#). Practically all the [furfural](#) now prepared technically is from oat hulls.<sup>2</sup>

This preparation is referenced from:

- [Org. Syn. Coll. Vol. 1, 276](#)
- [Org. Syn. Coll. Vol. 1, 283](#)
- [Org. Syn. Coll. Vol. 2, 302](#)

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### References and Notes

1. Stone, Am. Chem. J. **13**, 73 (1891); J. Anal. Appl. Chem. **5**, 421 (1891); Ber. **24**, 3019 (1891).
  2. The literature on the preparation of furfural is too extensive for citation here. This literature has been reviewed in "Furfural and Its Derivatives," Bulletin 2 (1928), The Miner Laboratories, Chicago, Ill. Some additional references are: Heuser and Scherer, Brennstoff-Chem. **4**, 97 (1923); Gierisch, Cellulosechem. **6**, 61 (1925); Scurti, Atti II congresso naz. chim. pura applicata 317 (1926) [C. A. **22**, 4665 (1928)]; Vickers Ltd. and Lucas, Brit. pat. 298,800 [C. A. **23**, 3098 (1929)]; Brownlee, Ind. Eng. Chem. **19**, 422 (1927); Clemen, Agr. Expt. Sta. Circ. **339**, 1 (1928) [C. A. **23**, 1969 (1929)]; Oborin Zhur. Prikladnoi Khimii **3**, 231 (1930) [C. A. **24**, 4029 (1930)]; Klingstedt, Zellstoff u. Papier **8**, 471 (1928) [C. A. **24**, 4784 (1930)].
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**Appendix**  
**Chemical Abstracts Nomenclature (Collective Index Number);**  
**(Registry Number)**

aldehyde

pentoses

hydrochloric or sulfuric acids

sulfuric acid (7664-93-9)

hydrochloric acid (7647-01-0)

sodium hydroxide (1310-73-2)

chloroform (67-66-3)

Furfural,  
2-Furaldehyde (98-01-1)